

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. N,N-Dimethylmethanamine, (Trimethylamine); C ₃ H ₉ N; [75-50-3] 2. Methanol; CH ₄ O; [67-56-1]		Halban, H. Z. Phys. Chem. <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C ₃ H ₉ N in solution/ mol dm ⁻³	p _{C₃H₉N} /mmHg	Concentration of C ₃ H ₉ N in solution/ concentration in gas phase	Mole fraction in solution* x _{C₃H₉N}
298.2	0.368	9.63	712	0.0148
	0.510	13.2	718	0.0204
	0.511	13.5	704	0.0204
<p>* Calculated by the compiler, on the assumption that dissolution of gas caused negligible change of volume of the liquid phase, using the density of methanol given in ref. (1).</p>				
AUXILIARY INFORMATION				
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
<p>The partial pressures of trimethylamine above solutions of concentrations determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.</p>		1. Hydrated chloride from Kahlbaum reacted with KOH; gas dried with NaOH. 2. Kahlbaum <i>methylalkohol I</i> distilled from CaO. b.p. 64.3-65.3°C (752 mmHg).		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. Timmermans, J. <i>Physico-Chemical Constants of Pure Organic Compounds</i> Vol. 2, Elsevier, Amsterdam. <u>1965</u> 2. Gaus, Z. <i>Anorg. Chem.</i> <u>1900</u> , 25, 236. 3. Abegg, R.; Riesenfeld, H. Z. <i>Phys. Chem.</i> <u>1902</u> , 40, 84.		

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. N,N-Dimethylmethanamine, (Trimethylamine); C ₃ H ₉ N; [75-50-3] 2. Ethanol; C ₂ H ₆ O; [557-75-5]		Halban, H. Z. Phys. Chem. <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C ₃ H ₉ N in solution/ mol dm ³	p _{C₃H₉N} /mmHg	Concentration of C ₃ H ₉ N in solution/ concentration in gas phase	Mole fraction in solution* x _{C₃H₉N}
298.2	0.1914	7.62	467	0.0111
	0.299	11.8	472	0.0172
	0.299	11.7	474	0.0172
<p>* Calculated by the compiler using the density of the solvent given in ref. (1), on the assumption that dissolution of gas caused negligible change of volume of the liquid phase.</p>				
AUXILIARY INFORMATION				
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
<p>The partial pressures of trimethylamine above solutions of concentrations determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.</p>		1. Hydrated chloride from Kahlbaum reacted with KOH; gas dried with NaOH. 2. Dried over CaO and distilled; b.p. 78.1°C (760 mmHg).		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. Timmermans, J. <i>Physico-Chemical Constants of Pure Organic Compounds</i> Vol. 2, Elsevier, Amsterdam. <u>1965</u> 2. Gaus. Z. Anorg. Chem. <u>1900</u> , 25, 236. 3. Abegg, R; Riesenfeld, H. Z. Phys. Chem. <u>1902</u> , 40, 84.		

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. N,N-Dimethylmethanamine, (Trimethylamine); C ₃ H ₉ N; [75-50-3] 2. 1-Propanol; C ₃ H ₈ O; [71-23-8]		Halban, H. <i>Z. Phys. Chem.</i> <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C ₃ H ₉ N in solution/ mol dm ⁻³	p _{C₃H₉N} /mmHg	Concentration of C ₃ H ₉ N in solution/ concentration in gas phase	Mole fraction in solution* x _{C₃H₉N}
298.2	0.219	8.55	482	0.0162
	0.329	13.1	467	0.0241
	0.330	13.2	467	0.0242
<p>* Calculated by the compiler, using the density of the solvent given in ref. (1), on the assumption that dissolution of gas caused negligible change of volume of the liquid phase.</p>				
AUXILIARY INFORMATION				
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
<p>The partial pressures of trimethylamine above solutions of concentrations determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.</p>		1. Hydrated chloride from Kahlbaum; reacted with KOH; gas dried with NaOH. 2. Dried over KOH and distilled; b.p. 96.8-97.1°C (754 mmHg).		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. Timmermans, J. <i>Physico-Chemical Constants of Pure Organic Compounds</i> Vol. 2, Elsevier, Amsterdam. 1965 2. Gaus, <i>Z. Anorg. Chem.</i> 1900, 25, 236. 3. Abegg, R.; Riesenfeld, H. <i>Z. Phys. Chem.</i> 1902, 40, 84.		

COMPONENTS: 1. N,N-Dimethylmethanamine (trimethylamine); C ₃ H ₉ N; [75-50-3] 2. 1-Butanol; C ₄ H ₁₀ O; [71-36-3]	ORIGINAL MEASUREMENTS: Gerrard, W. <i>Solubility of Gases and Liquids, Plenum 1976</i> , Chapter 10.																
VARIABLES: Temperature, pressure	PREPARED BY: C. L. Young																
EXPERIMENTAL VALUES: <table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th style="text-align: left;">T/K</th> <th style="text-align: left;">P/mmHg</th> <th style="text-align: left;">P/10⁵Pa</th> <th style="text-align: left;">Mole fraction of trimethylamine in liquid, $x_{(CH_3)_3N}$</th> </tr> </thead> <tbody> <tr> <td>293.15</td> <td>760</td> <td>1.013</td> <td>0.571</td> </tr> <tr> <td>298.15</td> <td>700</td> <td>0.933</td> <td>0.488</td> </tr> <tr> <td></td> <td>760</td> <td>1.013</td> <td>0.510</td> </tr> </tbody> </table>		T/K	P/mmHg	P/10 ⁵ Pa	Mole fraction of trimethylamine in liquid, $x_{(CH_3)_3N}$	293.15	760	1.013	0.571	298.15	700	0.933	0.488		760	1.013	0.510
T/K	P/mmHg	P/10 ⁵ Pa	Mole fraction of trimethylamine in liquid, $x_{(CH_3)_3N}$														
293.15	760	1.013	0.571														
298.15	700	0.933	0.488														
	760	1.013	0.510														
AUXILIARY INFORMATION																	
METHOD/APPARATUS/PROCEDURE: Amine was passed into a known weight of pure liquid in a bubbler tube at a total pressure measured by a manometer assembly. The amount of absorbed gas was estimated by weighing. The temperature was manually controlled to within 0.2K. The apparatus and procedure are described by Gerrard [1,2].	SOURCE AND PURITY OF MATERIALS: 1. British Drug Houses or Cambrian Gases sample. 2. Purified and attested by conventional procedures. ESTIMATED ERROR: $\delta T/K = \pm 0.1$; $\delta x/x = \pm 3\%$ (estimated by compiler) REFERENCES: 1. Gerrard, W. <i>J. Appl. Chem. Biotechnol.</i> <u>1972</u> , 22 623-650. 2. Gerrard, W. <i>Solubility of Gases and Liquids. Plenum Press, New York. 1976.</i> Chapter 1.																

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. N,N-Dimethylmethanamine, (Trimethylamine); C ₃ H ₉ N; [75-50-3] 2. 3-Methyl-1-Butanol; C ₅ H ₁₂ O; [123-51-3]		Halban, H. Z. Phys. Chem. <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C ₃ H ₉ N in solution/ mol dm ⁻³	p C ₃ H ₉ N/mmHg	Concentration of C ₃ H ₉ N in solution/ concentration in gas phase	Mole fraction in solution* x C ₃ H ₉ N
298.2	0.342	16.3	390	0.0358
	0.346	16.9	382	0.0362
	0.347	16.9	383	0.0363
<p>* Calculated by the compiler, using the density of the solvent given in ref. (1), on the assumption that dissolution of gas caused negligible change of volume of the liquid phase.</p>				
AUXILIARY INFORMATION				
METHOD APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
<p>The partial pressures of trimethylamine above solutions of concentrations determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.</p>		1. Hydrated chloride from Kahlbaum; reacted with KOH; gas dried with NaOH. 2. Kahlbaum <i>Amylalkohol frei von Pyridin</i> distilled twice from Na; fraction distilling 130.6-131.0°C (752 mmHg) was collected.		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. <i>Lange's Handbook of Chemistry</i> (12th edition), McGraw-Hill, New York. 1979. 2. Gaus, Z. <i>Anorg. Chem.</i> 1900, 25, 236. 3. Abegg, R.; Riesenfeld, H. Z. <i>Phys. Chem.</i> 1902, 40, 84.		

EXPERIMENTAL VALUES:			Mole fraction of trimethylamine in liquid, x (CH ₃) ₃ N
T/K	P/mmHg	P/10 ⁵ Pa	
278.15	100	0.133	0.263
	200	0.267	0.401
	300	0.400	0.510
	400	0.533	0.607
	500	0.667	0.691
	600	0.800	0.790
	700	0.933	0.882
	760	1.013	0.936
283.15	100	0.133	0.240
	200	0.267	0.392
	300	0.400	0.472
	400	0.533	0.538
	500	0.667	0.602
	600	0.800	0.675
	700	0.933	0.752
	760	1.013	0.800
293.15	760	1.013	0.636
298.15	100	0.133	0.157
	200	0.267	0.251
	300	0.400	0.320
	400	0.533	0.382
	500	0.667	0.436
	600	0.800	0.486
	700	0.933	0.538
	760	1.013	0.568

AUXILIARY INFORMATION

METHOD/APPARATUS/PROCEDURE:

Amine was passed into a known weight of pure liquid in a bubbler tube at a total pressure measured by a manometer assembly. The amount of absorbed gas was estimated by weighing. The temperature was manually controlled to within 0.2K. The apparatus and procedure are described by Gerrard [1,2].

SOURCE AND PURITY OF MATERIALS:

1. British Drug Houses or Cambrian Gases sample.
2. Purified and attested by conventional procedures.

ESTIMATED ERROR:

$\delta T/K = \pm 0.1$; $\delta x/x = \pm 3\%$
(estimated by compiler)

REFERENCES:

1. Gerrard, W. *J. Appl. Chem. Biotechnol.* **1972**, *22* 623-650.
2. Gerrard, W. *Solubility of Gases and Liquids*. Plenum Press, New York. **1976**. Chapter 1.

COMPONENTS:

1. N,N-Dimethylmethanamine, (Trimethylamine); C₃H₉N; [75-50-3]
2. 1-Octanol; C₈H₁₈O; [111-87-5]

ORIGINAL MEASUREMENTS:

Gerrard, W.
Solubility of Gases and Liquids,
Plenum **1976**, Chapter 10.

VARIABLES:

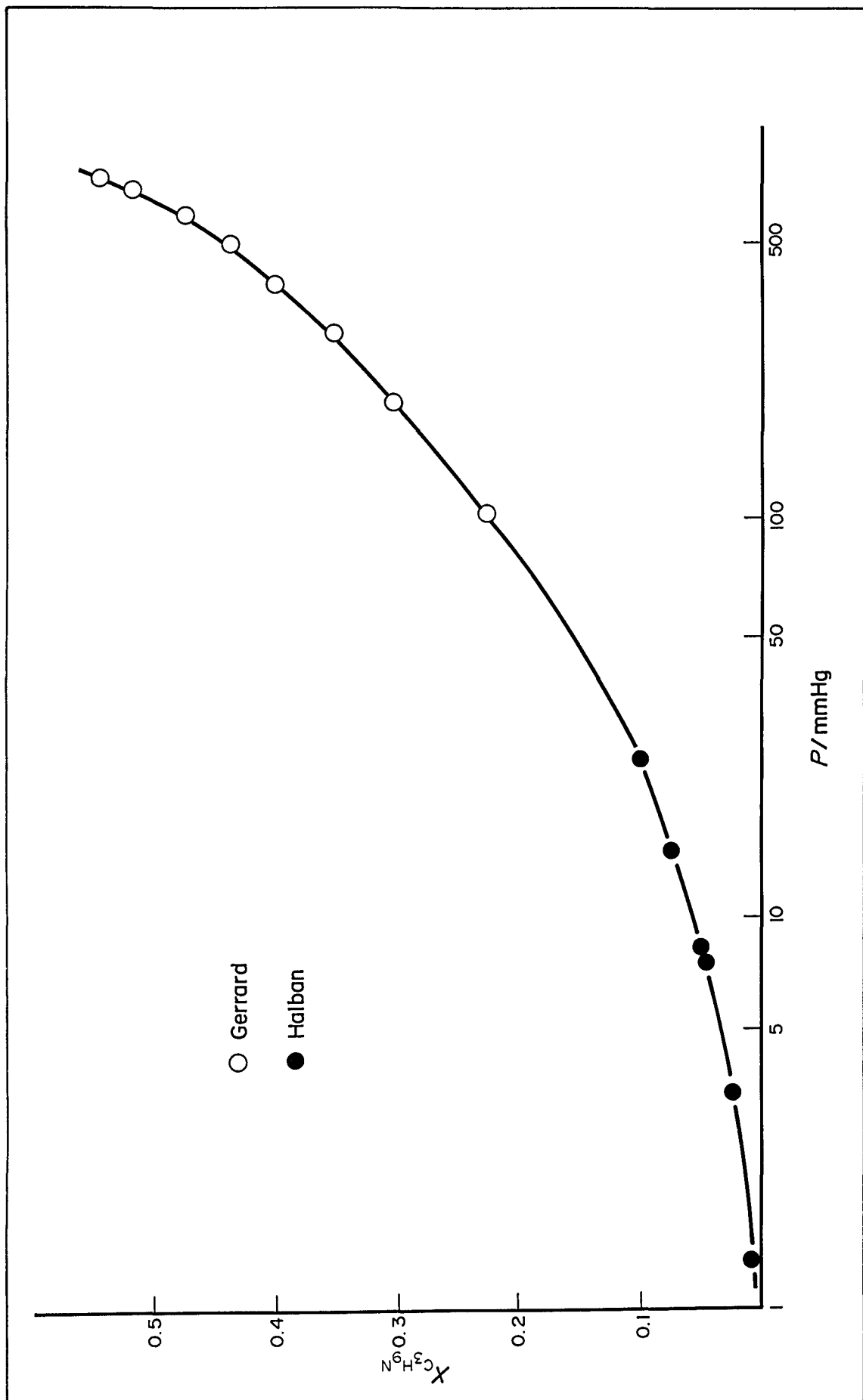
Temperature, pressure

PREPARED BY:

C. L. Young

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. <i>N,N</i> -Dimethylmethanamine, (Trimethylamine); C_3H_9N ; [75-50-3] 2. Benzenemethanol (benzyl alcohol); C_7H_8O ; [100-51-6]		Halban, H. <i>Z. Phys. Chem.</i> <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C_3H_9N in solution/ mol dm ³	$p_{C_3H_9N}$ /mmHg	Concentration of C_3H_9N in solution/ concentration in gas phase	Mole fraction in solution* $x_{C_3H_9N}$
298.2	0.0950	1.35	1308	0.0098
	0.247	3.57	1282	0.0250
	0.447	7.39	1127	0.0443
	0.486	8.37	1081	0.0480
	0.763	14.6	973	0.0734
	0.769	14.7	968	0.0739
	1.040	25.0	775	0.0975
* Calculated by the compiler, using the density of the solvent given in ref. (1), on the assumption that dissolution of gas caused negligible change of volume of the liquid phase.				
AUXILIARY INFORMATION				
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
The partial pressures of trimethyl- amine above solutions of concentrat- ions determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/ oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.		1. Hydrated chloride from Kahlbaum; reacted with KOH; gas dried with NaOH. 2. From Schmitz, Düsseldorf; b.p. 204.8-205.2°C.		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. Dreisbach, R.R. <i>Physical Properties of Chemical Compounds</i> , Vol. 1, A.C.S. Washington. <u>1955</u> . 2. Gaus, <i>Z. Anorg. Chem.</i> <u>1900</u> , 25, 236. 3. Abegg, R.; Riesenfeld, H. <i>Z. Phys. Chem.</i> <u>1902</u> , 40, 84.		

COMPONENTS:			ORIGINAL MEASUREMENTS:
1. N,N-Dimethylmethanamine (trimethylamine); C ₃ H ₉ N; [75-50-3]			Gerrard, W.
2. Benzenemethanol (Benzyl alcohol); C ₇ H ₈ O; [100-51-6]			<i>Solubility of Gases and Liquids</i> , Plenum <u>1976</u> , Chapter 10.
VARIABLES:			PREPARED BY:
Temperature, pressure			C. L. Young
EXPERIMENTAL VALUES:			
T/K	P/mmHg	P/10 ⁵ Pa	Mole fraction of trimethylamine in liquid, x (CH ₃) ₃ N
278.15	100	0.133	0.315
	200	0.267	0.407
	300	0.400	0.484
	400	0.533	0.557
	500	0.667	0.644
	600	0.800	0.742
	700	0.933	0.866
	760	1.013	0.925
283.15	100	0.133	0.293
	200	0.267	0.372
	300	0.400	0.448
	400	0.533	0.508
	500	0.667	0.574
	600	0.800	0.644
	700	0.933	0.725
	760	1.013	0.776
293.15	760	1.013	0.585
298.15	100	0.133	0.224
	200	0.267	0.301
	300	0.400	0.353
	400	0.533	0.395
	500	0.667	0.436
	600	0.800	0.475
	700	0.933	0.516
	760	1.013	0.541
AUXILIARY INFORMATION			
METHOD/APPARATUS/PROCEDURE:			SOURCE AND PURITY OF MATERIALS:
Amine was passed into a known weight of pure liquid in a bubbler tube at a total pressure measured by a manometer assembly. The amount of absorbed gas was estimated by weighing. The temperature was manually controlled to within 0.2K. The apparatus and procedure are described by Gerrard [1,2].			1. British Drug Houses or Cambrian Gases sample.
			2. Purified and attested by conventional procedures.
			ESTIMATED ERROR:
			$\delta T/K = \pm 0.1$; $\delta x/x = \pm 3\%$
			(estimated by compiler)
			REFERENCES:
			1. Gerrard, W. <i>J. Appl. Chem. Biotechnol.</i> <u>1972</u> , 22 523-650.
			2. Gerrard, W. <i>Solubility of Gases and Liquids</i> . Plenum Press, New York. <u>1976</u> . Chapter 1.



COMPONENTS:			ORIGINAL MEASUREMENTS:
1. N,N-Dimethylmethanamine (trimethylamine); C ₃ H ₉ N; [75-50-3]			Gerrard, W.
2. 1,2-Ethanediol (Ethylene glycol); C ₂ H ₆ O ₂ ; [107-21-1]			<i>Solubility of Gases and Liquids</i> , <i>Plenum 1976</i> , Chapter 10.
VARIABLES:			PREPARED BY:
Temperature, pressure			C. L. Young
EXPERIMENTAL VALUES:			
			Mole fraction of trimethylamine in liquid, x (CH ₃) ₂ N
T/K	P/mmHg	P/10 ⁵ Pa	
283.15	100	0.133	0.116
	200	0.267	0.216
	300	0.400	0.304
	400	0.533	0.384
	500	0.667	0.455
	600	0.800	0.531
	700	0.933	0.616
	760	1.013	0.672
298.15	100	0.133	0.058
	200	0.267	0.120
	300	0.400	0.180
	400	0.533	0.232
	500	0.667	0.285
	600	0.800	0.331
	700	0.933	0.378
	760	1.013	0.396
AUXILIARY INFORMATION			
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:	
<p>Amine was passed into a known weight of pure liquid in a bubbler tube at a total pressure measured by a manometer assembly. The amount of absorbed gas was estimated by weighing. The temperature was manually controlled to within 0.2K.</p> <p>The apparatus and procedure are described by Gerrard [1,2].</p>		1. British Drug Houses or Cambrian Gases sample. 2. Purified and attested by conventional procedures.	
		ESTIMATED ERROR:	
		$\delta T/K = \pm 0.1$; $\delta x/x = \pm 3\%$ (estimated by compiler)	
		REFERENCES:	
		1. Gerrard, W. <i>J. Appl. Chem. Biotechnol.</i> <u>1972</u> , 22 623-650.	
		2. Gerrard, W. <i>Solubility of Gases and Liquids</i> . <i>Plenum Press, New York.</i> <u>1976</u> . Chapter 1.	

COMPONENTS:		ORIGINAL MEASUREMENTS:		
1. N,N-Dimethylmethanamine, (Trimethylamine); C ₃ H ₉ N; [75-50-3] 2. 2-Propanone, (acetone); C ₃ H ₆ O; [67-64-1]		Halban, H. <i>Z. Phys. Chem.</i> <u>1913</u> , 84, 129-159.		
VARIABLES:		PREPARED BY:		
Concentration		P. G. T. Fogg		
EXPERIMENTAL VALUES:				
T/K	Concentration of C ₃ H ₉ N in solution/ mol dm ⁻³	<i>p</i> C ₃ H ₉ N/mmHg	Concentration of C ₃ H ₉ N in solution/ concentration in gas phase	Mole fraction in solution* <i>x</i> C ₃ H ₉ N
298.2	0.212	51.8	76.2	0.0153
	0.237	57.8	76.7	0.0171
	0.281	68.5	76.2	0.0202
	0.325	77.6	77.7	0.0233
* Calculated by the compiler, using the density of the solvent given in ref. (1), on the assumption that dissolution of gas caused negligible change of volume of the liquid phase.				
AUXILIARY INFORMATION				
METHOD/APPARATUS/PROCEDURE:		SOURCE AND PURITY OF MATERIALS:		
<p>The partial pressures of trimethylamine above solutions of concentrations determined by titration, were measured by a dynamic method (refs. (2) & (3)). Mixtures of hydrogen and oxygen, produced by electrolysis of sodium hydroxide solution, were passed through each solution of trimethylamine. The trimethylamine in the gas stream was absorbed in hydrochloric acid and estimated from changes in electrical conductivity due to partial neutralisation of the acid. The volumes of hydrogen/oxygen gas mixture produced by electrolysis were found from the barometric pressure and changes in a copper voltameter in series with the cell for producing the gas. The partial pressures of trimethylamine were calculated on the assumption that equilibrium was established between trimethylamine in solution and that in the gas phase, during passage of hydrogen/oxygen mixture through the solution.</p>		1. Hydrated chloride from Kahlbaum; reacted with KOH; gas dried with NaOH. 2. Kahlbaum <i>aus der Bisulfitverbindung</i> ; dried over anhydrous CuSO ₄ ; b.p. 56.4°C (769 mmHg).		
		ESTIMATED ERROR:		
		REFERENCES:		
		1. <i>Lange's Handbook of Chemistry</i> (12th edition), McGraw-Hill, New York, 1979. 2. Gaus, <i>Z. Anorg. Chem.</i> 1900, 25, 236. 3. Abegg, R.; Riesenfeld, H. <i>Z. Phys. Chem.</i> 1902, 40, 84.		

COMPONENTS: 1. <i>N,N</i> -Dimethylmethanamine (trimethylamine); C ₃ H ₉ N; [75-50-3] 2. 1,2,3-Propanetriol (Glycerol); C ₃ H ₈ O ₃ ; [56-81-5]			ORIGINAL MEASUREMENTS: Gerrard, W. <i>Solubility of Gases and Liquids</i> , Plenum 1976, Chapter 10.
VARIABLES: Temperature, pressure			PREPARED BY: C. L. Young
EXPERIMENTAL VALUES:			
			Mole fraction of trimethylamine in liquid, x (CH ₃) ₃ N
T/K	P/mmHg	P/10 ⁵ Pa	
283.15	100	0.133	0.134
	200	0.267	0.238
	300	0.400	0.316
	400	0.533	0.395
	500	0.667	0.464
	600	0.800	0.530
	700	0.933	0.604
	760	1.013	0.657
298.15	100	0.133	0.075
	200	0.267	0.129
	300	0.400	0.177
	400	0.533	0.224
	500	0.667	0.272
	600	0.800	0.322
	700	0.933	0.376
	760	1.013	0.408
AUXILIARY INFORMATION			
METHOD/APPARATUS/PROCEDURE: Amine was passed into a known weight of pure liquid in a bubbler tube at a total pressure measured by a manometer assembly. The amount of absorbed gas was estimated by weighing. The temperature was manually controlled to within 0.2K. The apparatus and procedure are described by Gerrard [1,2].		SOURCE AND PURITY OF MATERIALS: 1. British Drug Houses or Cambrian Gases sample. 2. Purified and attested by conventional procedures.	
		ESTIMATED ERROR: $\delta T/K = \pm 0.1$; $\delta x/x = \pm 3\%$ (estimated by compiler)	
		REFERENCES: 1. Gerrard, W. <i>J. Appl. Chem. Biotechnol.</i> 1972, 22 623-650. 2. Gerrard, W. <i>Solubility of Gases and Liquids</i> . Plenum Press, New York. 1976. Chapter 1.	