

<b>COMPONENTS:</b> (1) Dysprosium bromide; DyBr <sub>3</sub> ; [14456-48-5] (2) 1,2-Diethoxyethane; C <sub>6</sub> H <sub>14</sub> O <sub>2</sub> ; [629-14-1]	<b>ORIGINAL MEASUREMENTS:</b> Kirmse, E.M. <i>Tr. II Vses. Konf. po Teor. Rastvorov</i> <u>1971</u> , 200-6.
<b>VARIABLES:</b> T/K = 298	<b>PREPARED BY:</b> T. Mioduski
<b>EXPERIMENTAL VALUES:</b> <p>The solubility of DyBr<sub>3</sub> in 1,2-diethoxyethane at 25°C was given as</p> <p style="text-align: center;">1.2 mass %</p> <p>The corresponding value of the molality calculated by the compiler is</p> <p style="text-align: center;">0.0302 mol kg<sup>-1</sup></p> <p>The nature of the solid phase was not specified.</p>	
<b>AUXILIARY INFORMATION</b>	
<b>METHOD/APPARATUS/PROCEDURE:</b> Experimental details not given, but were probably similar to previous works of the author which are compiled throughout this volume.	<b>SOURCE AND PURITY OF MATERIALS:</b> Nothing specified, but based on previous work by the author, the anhydrous salt was probably prepared by the method of Taylor and Carter (1). <b>ESTIMATED ERROR:</b> Nothing specified. <b>REFERENCES:</b> 1. Taylor, M.D.; Carter, C.P. <i>J. Inorg. Nucl. Chem.</i> <u>1962</u> , 24, 387.

<b>COMPONENTS:</b> (1) Dysprosium bromide; DyBr <sub>3</sub> ; [14456-48-5] (2) Tetrahydrofuran; C <sub>4</sub> H <sub>8</sub> O; [109-99-9]	<b>ORIGINAL MEASUREMENTS:</b> Rossmannith, K. <i>Monatsh. Chem.</i> <u>1966</u> , 97, 1357-64.
<b>VARIABLES:</b> Room Temperature: T/K = 294-296	<b>PREPARED BY:</b> T. Mioduski
<b>EXPERIMENTAL VALUES:</b> <p>The solubility of DyBr<sub>3</sub> in tetrahydrofuran at 21-23°C was reported to be 0.26 g per 100 ml of solution (0.0065 mol dm<sup>-3</sup>, compiler).</p>	
<b>AUXILIARY INFORMATION</b>	
<b>METHOD/APPARATUS/PROCEDURE:</b> Isothermal method employed. The solution was equilibrated in an extractor with agitation for 60-80 hours at room temperature. Dysprosium was determined by the oxalate method and by titration with EDTA using Xylenol Orange indicator. The solvent was determined by difference. Anhydrous materials were handled in a dry box through which was passed a stream of nitrogen free of carbon dioxide. The solid phase is DyBr <sub>3</sub> ·3.5C <sub>4</sub> H <sub>8</sub> O.	<b>SOURCE AND PURITY OF MATERIALS:</b> Sources and purities of initial materials not specified. DyBr <sub>3</sub> was prepared by conversion of the oxide by high temperature reaction with an excess of NH <sub>4</sub> Br followed by heating the product in a stream of dry nitrogen, and then in vacuum to remove unreacted NH <sub>4</sub> Br. Tetrahydrofuran was distilled from LiAlH <sub>4</sub> . <b>ESTIMATED ERROR:</b> Nothing specified.
	<b>REFERENCES:</b>

<b>COMPONENTS:</b> (1) Dysprosium bromide; DyBr <sub>3</sub> ; [14456-48-5] (2) 1,4-Dioxane; C <sub>4</sub> H <sub>8</sub> O <sub>2</sub> ; [123-91-1]	<b>ORIGINAL MEASUREMENTS:</b> Kirmse, E.M.; Zwietasch, K.J.; Tirschmann, J.; Oelsner, L.; Niedergeases, U. <i>Z. Chem.</i> <u>1968</u> , <i>8</i> , 472-3. Kirmse, E.M. <i>Tr. II Vses. Konf. po Teor. Rastvorov</i> <u>1971</u> , 200-6.
<b>VARIABLES:</b> Room Temperature: T/K around 298	<b>PREPARED BY:</b> T. Mioduski
<b>EXPERIMENTAL VALUES:</b> <p>The solubility of DyBr<sub>3</sub> in p-dioxane at about 25°C was given as</p> <p style="text-align: center;">0.9<sub>5</sub> mass %</p> <p>the corresponding molality calculated by the compiler is</p> <p style="text-align: center;">0.0238 mol kg<sup>-1</sup></p> <p>The nature of the solid phase was not specified.</p>	
<b>AUXILIARY INFORMATION</b>	
<b>METHOD/APPARATUS/PROCEDURE:</b> The solute-solvent mixtures were isothermally agitated at 25°C or at room temperature. Authors state that the difference found for the solubility was within experimental error limits. Dy was determined by complexometric titration. No other details given.	<b>SOURCE AND PURITY OF MATERIALS:</b> The anhydrous salt was prepared by the method of Taylor and Carter (1). No other information given. <b>ESTIMATED ERROR:</b> Nothing specified. <b>REFERENCES:</b> 1. Taylor, M.D.; Carter, C.P. <i>J. Inorg. Nucl. Chem.</i> <u>1962</u> , <i>24</i> , 387.

<b>COMPONENTS:</b> (1) Dysprosium bromide; DyBr <sub>3</sub> ; [14456-48-5]  (2) 2-Propanamine; iso-C <sub>3</sub> H <sub>9</sub> N; [75-31-0]	<b>ORIGINAL MEASUREMENTS:</b> Kirmse, E.M.  <i>Tr. II Uses. Konf. po Teor. Rastvorov</i> <u>1971</u> , 200-6.
<b>VARIABLES:</b>  T/K = 298	<b>PREPARED BY:</b>  T. Mioduski
<b>EXPERIMENTAL VALUES:</b>  The solubility of DyBr <sub>3</sub> in 2-propanamine at 25°C was reported as $43.0 \text{ mass } \%$  The corresponding molality calculated by the compiler is $1.876 \text{ mol kg}^{-1}$  The nature of the solid phase was not specified.	
<b>AUXILIARY INFORMATION</b>	
<b>METHOD/APPARATUS/PROCEDURE:</b> Experimental details not given, but were probably similar to previous works of the author which are compiled throughout this volume.	<b>SOURCE AND PURITY OF MATERIALS:</b> Nothing specified, but based on previous work by the author, the anhydrous salt was probably prepared by the method of Taylor and Carter (1).  <b>ESTIMATED ERROR:</b>  Nothing specified.  <b>REFERENCES:</b> 1. Taylor, M.D.; Carter, C.P. <i>J. Inorg. Nucl. Chem.</i> <u>1962</u> , 24, 387.