COMPONENTS:	ORIGINAL MEASUREMENTS:
(1) Ytterbium iodide; YbI_3; [13813-44-0] (2) 1,4-Dioxane; $C_4H_8O_2$; [123-81-1]	Kirmse, E.M.; Zwietasch, K.J.; Tirschmann, J.; Oelsner, L.; Niedergesaess, U. *Z. Chem.* 1968, *8*, 472-3. Kirmse, E.M. *Tr. II Vses. Konf. po Teor. Rastvorov.* 1971, 200-6.
VARIABLES: Room Temperature: T/K around 298	**PREPARED BY:** T. Mioduski

EXPERIMENTAL VALUES:

The solubility of YbI_3 in p-dioxane at around 25°C was given as

0.1_5 mass %

The corresponding molality calculated by the compiler is

0.0027 mol kg^{-1}

The nature of the solid phase was not specified.

AUXILIARY INFORMATION

METHOD/APPARATUS/PROCEDURE:	SOURCE AND PURITY OF MATERIALS:
The solute-solvent mixtures were isothermally agitated at 25°C or at room temperature. Authors state that the difference found for the solubility was within experimental error limits. Yb was determined by complexometric titration. No other details given.	The anhydrous salt was prepared by the method of Taylor and Carter (1). No other information given.
	ESTIMATED ERROR: Nothing specified.
	REFERENCES: 1. Taylor, M.D.; Carter, C.P. *J. Inorg. Nucl. Chem.* 1962, *24*, 387.

COMPONENTS:	ORIGINAL MEASUREMENTS:
(1) Ytterbium iodide; YbI_3; [13813-44-0] (2) Alkyl amines	Kirmse, E.M. *Tr. II Vses. Konf. po Teor. Rastvorov* 1971, 200-6.
VARIABLES: T/K = 298	**PREPARED BY:** T. Mioduski and M. Salomon

EXPERIMENTAL VALUES:

solvent			YbI_3 solubility[a] mass %	mol kg^{-1}
1-propanamine;	C_3H_9N;	[107-10-8]	1.7	0.031
2-propanamine;	iso-C_3H_9N;	[75-31-0]	1.2	0.022
1-butanamine;	n-$C_4H_{11}N$;	[109-73-9]	0.4	0.007
2-butanamine;	iso-$C_4H_{11}N$;	[13952-84-6]	8.9	0.176

[a]Molalities calculated by the compilers.

AUXILIARY INFORMATION

METHOD/APPARATUS/PROCEDURE:
Experimental details not given, but were probably similar to previous works of the author which are compiled throughout this volume.

Nature of solid phases not specified.

SOURCE AND PURITY OF MATERIALS:
Nothing specified, but based on previous work by the author, the anhydrous salt was probably prepared by the method of Taylor and Carter (1).

ESTIMATED ERROR:
Nothing specified.

REFERENCES:
1. Taylor, M.D.; Carter, C.P.
J. Inorg. Nucl. Chem. 1962, *24*, 387.