

<b>COMPONENTS:</b> (1) Lanthanum iodide; $\text{LaI}_3$ ; [13813-22-4]  (2) N,N-Dimethylformamide; $\text{C}_3\text{H}_7\text{NO}$ ; [68-12-2]	<b>ORIGINAL MEASUREMENTS:</b> Moeller, T.; Galasyn, V.  <i>J. Inorg. Nucl. Chem.</i> <u>1960</u> , <i>12</i> , 259-65.
<b>VARIABLES:</b>  T/K = 298.15	<b>PREPARED BY:</b>  M. Salomon
<b>EXPERIMENTAL VALUES:</b>  The solubility of $\text{LaI}_3$ in $\text{HCON}(\text{CH}_3)_2$ at $25^\circ\text{C}$ was reported as $578.7 \text{ g dm}^{-3}$ and as $0.5240 \text{ mol dm}^{-3}$  The solid phase is the solvate $\text{LaI}_3 \cdot 8\text{HCON}(\text{CH}_3)_2$ . The melting point (sealed tube method) of this solvate given as $73.0\text{--}75.0^\circ\text{C}$ .	
<b>AUXILIARY INFORMATION</b>	
<b>METHOD/APPARATUS/PROCEDURE:</b> Authors state that solubilities were determined by analysis of aliquots after equilibration at $25 \pm 0.025^\circ\text{C}$ , and that techniques were generally similar to those described in (1).  The rare earth content was determined by complexometric titration with EDTA at $60^\circ\text{C}$ . Iodide was determined by the Volhard method, and carbon, hydrogen, and nitrogen by usual microanalytical techniques.	<b>SOURCE AND PURITY OF MATERIALS:</b> The initial material was the rare earth oxide of 99.9+% purity. Iodides were prepd by two methods. 1. Acetyl iodide method (2) where the hydrated acetate is treated with acetyl iodide in benzene. Acetyl iodide prepd as in (3). 2. The iodide was prepd by metathesis by reaction of the hydrated $\text{LaCl}_3$ with KI in DMF followed by addition of benzene and distillation of the benzene-water azeotrope.  For both preparations the solvate $\text{LaI}_3 \cdot 8\text{DMF}$ was recrystallized from DMF by addition of ether.  The solvent, DMF, was prepared as in (4,5), and its electrolytic conductance was $\geq 3.7 \times 10^{-7} \text{ S cm}^{-1}$ at $25^\circ\text{C}$ .
<b>REFERENCES:</b> 1. Moeller, T.; Cullen, G.W. <i>J. Inorg. Nucl. Chem.</i> <u>1959</u> , <i>10</i> , 148. 2. Watt, G.W.; Gentile, P.S.; Helvenston, E.P. <i>J. Am. Chem. Soc.</i> <u>1955</u> , <i>77</i> , 2752. 3. Biltz, H.; Biltz, W. <i>Laboratory Methods of Inorganic Chemistry (2nd Edition)</i> . John Wiley. N.Y. <u>1928</u> . 4. Leader, G.R.; Gormley, J.F. <i>J. Am. Chem. Soc.</i> <u>1951</u> , <i>73</i> , 5731. 5. Thomas, A.B.; Rochow, E.G. <i>J. Am. Chem. Soc.</i> <u>1957</u> , <i>79</i> , 1843.	<b>ESTIMATED ERROR:</b> Soly: precision around $\pm 0.1\%$ (compiler).  Temp: precision $\pm 0.025 \text{ K}$ (authors).