

<b>COMPONENTS:</b> 1. Beryllium selenite; $\text{BeSeO}_3$ ; [38333-62-9] 2. Water; $\text{H}_2\text{O}$ ; [7732-18-5]	<b>ORIGINAL MEASUREMENTS:</b> Ripan, R.; Vericeanu, G.  <i>Studia Univ. Babeş-Bolyai, Ser. Chim.</i> <u>1968, 13, 31-37</u>																				
<b>VARIABLES:</b> One temperature: 291 K	<b>PREPARED BY:</b> Mary R. Masson																				
<b>EXPERIMENTAL VALUES:</b> All concentrations are expressed in units of $\text{mol dm}^{-3}$ <table data-bbox="260 531 1042 756" style="margin-left: auto; margin-right: auto;"> <thead> <tr> <th>Concentration</th> <th><math>K_{\text{SO}}</math> <math>\text{mol}^2\text{dm}^6</math></th> <th>Mean <math>K_{\text{SO}}</math></th> <th><math>\text{p}K_{\text{SO}}</math></th> </tr> </thead> <tbody> <tr> <td><math>8.501 \times 10^{-5}</math></td> <td><math>7.2 \times 10^{-9}</math></td> <td rowspan="2"><math>1.1 \pm 0.5 \times 10^{-8}</math></td> <td rowspan="2">7.96</td> </tr> <tr> <td><math>8.451 \times 10^{-5}</math></td> <td><math>7.1 \times 10^{-9}</math></td> </tr> <tr> <td><math>9.150 \times 10^{-5}</math></td> <td><math>8.4 \times 10^{-9}</math></td> <td rowspan="4" style="text-align: center;"><math>\text{mol}^2\text{dm}^{-6}</math></td> <td></td> </tr> <tr> <td><math>1.141 \times 10^{-4}</math></td> <td><math>1.3 \times 10^{-8}</math></td> </tr> <tr> <td><math>1.186 \times 10^{-4}</math></td> <td><math>1.4 \times 10^{-8}</math></td> </tr> <tr> <td><math>1.189 \times 10^{-4}</math></td> <td><math>1.4 \times 10^{-8}</math></td> </tr> </tbody> </table> <p>The concentration <math>c</math> in the saturated solution was calculated from the measured conductivity <math>\kappa</math> from the equation</p> $c = \frac{1000\kappa}{\Lambda^\circ}$ <p><u>Compiler's note</u>            Neither in the determination of the ionic conductivity of the selenite ion nor in the evaluation of the solubility product was hydrolysis of the selenite ion taken into account. This would give rise to errors, since, for example, in a 0.001M solution, <math>[\text{SeO}_3^{2-}] = 0.000955M</math>, <math>[\text{HSeO}_3^-] = 0.000045M</math> and <math>[\text{OH}^-] = 0.000045M</math>, and hydroxide and hydrogen selenite have different ionic conductivities from selenite. If the ionic conductivity of hydrogen selenite were known, the experimental results could have been interpreted correctly (cf. ref. 2), but this value does not seem to be available. However, because the calibration and sample solutions had concentrations of about the same order of magnitude, the errors would cancel to some extent, but the <math>K_{\text{SO}}</math> value cannot be regarded as reliable.</p>		Concentration	$K_{\text{SO}}$ $\text{mol}^2\text{dm}^6$	Mean $K_{\text{SO}}$	$\text{p}K_{\text{SO}}$	$8.501 \times 10^{-5}$	$7.2 \times 10^{-9}$	$1.1 \pm 0.5 \times 10^{-8}$	7.96	$8.451 \times 10^{-5}$	$7.1 \times 10^{-9}$	$9.150 \times 10^{-5}$	$8.4 \times 10^{-9}$	$\text{mol}^2\text{dm}^{-6}$		$1.141 \times 10^{-4}$	$1.3 \times 10^{-8}$	$1.186 \times 10^{-4}$	$1.4 \times 10^{-8}$	$1.189 \times 10^{-4}$	$1.4 \times 10^{-8}$
Concentration	$K_{\text{SO}}$ $\text{mol}^2\text{dm}^6$	Mean $K_{\text{SO}}$	$\text{p}K_{\text{SO}}$																		
$8.501 \times 10^{-5}$	$7.2 \times 10^{-9}$	$1.1 \pm 0.5 \times 10^{-8}$	7.96																		
$8.451 \times 10^{-5}$	$7.1 \times 10^{-9}$																				
$9.150 \times 10^{-5}$	$8.4 \times 10^{-9}$	$\text{mol}^2\text{dm}^{-6}$																			
$1.141 \times 10^{-4}$	$1.3 \times 10^{-8}$																				
$1.186 \times 10^{-4}$	$1.4 \times 10^{-8}$																				
$1.189 \times 10^{-4}$	$1.4 \times 10^{-8}$																				
<b>AUXILIARY INFORMATION</b>																					
<b>METHOD APPARATUS/PROCEDURE:</b> The ionic conductivity of the selenite ion in water was determined by measuring the mean molar conductivities of a series of aqueous solutions of lithium selenite with concentrations below $2 \times 10^{-3} \text{mol dm}^{-3}$ . The mean molar conductivity at infinite dilution was found by extrapolation to be $172 \pm 2.9 \Omega^{-1}\text{cm}^2\text{mol}^{-1}$ at 18°C. At that temperature, the ionic conductivity of the lithium ion is $33.4 \Omega^{-1}\text{cm}^2\text{mol}^{-1}$ , so the ionic conductivity of the selenite ion is $105.2 \Omega^{-1}\text{cm}^2\text{mol}^{-1}$ . The well washed beryllium selenite was mixed with 150 ml of water and agitated intermittently for 1 hr. From this, 10-20 ml was taken in the conductivity cell and the conductivity measured. The sampling and measurement were repeated until a constant value for the conductivity was obtained. The value of $\Lambda^\circ = (\lambda_+ + \lambda_-)$ , the molar conductivity, was calculated with $\lambda_- = 105.2$ and $\lambda_+ = 74.8 \Omega^{-1}\text{cm}^2\text{mol}^{-1}$ .	<b>SOURCE AND PURITY OF MATERIALS:</b> Lithium selenite was prepared from sublimed selenium dioxide by neutralization with lithium hydroxide. An approximately 0.05M solution was standardized gravimetrically, and this was used to prepare the various dilutions. The composition of the selenite was checked by analysis.  <b>ESTIMATED ERROR:</b> Temperature: $\pm 0.5$ K Error in $K_{\text{SO}}$ (2s) = $0.5 \times 10^{-8}$ (compiler)  <b>REFERENCES:</b> 1. Landolt-Bornstein <i>Physikalisch-Chemische Tabellen II 1923</i> , p. 1105. 2. Monk, C.B. <i>J. Chem. Soc.</i> <u>1949</u> , 429.																				